**DEPARTMENT OF CHEMISTRY**

**Subject:** Inorganic Practical **Class:**  B.S (iii)/M. Sc. Previous

**Semester:** 1st **Test**  Final

**Session:** 2012 **Date:** 08- 01 -2013

**Maximum Marks** 50  **Time allowed:** 1

Name-------------------------------------- Father’s Name------------------------------------

Seat No -------------

**Note: Attempt all questions**

Q. No. 1 (a) Calculate the amount of given compounds for the preparation of 100mls of 0.05N solution

 (i) CaSO4 Atomic weight of Ca =40, S= 32, O = 16

1. H2SO4 Specific gravity = 1.8 and Percentage purity is 96
2. Calculate volume of 1M solution of HCl for preparation of its 100ml dilute solution of 0.1M solution

Q. No. 2 Define the following terms

 Chelates-----------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------

 Buffer soolution ---------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------

 Gravimetric Analysis-------------------------------------------------------------------------------------------------------------------------------------------------------------

 Titration--------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------

 Basicity and Acidity------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------

 Molarity---------------------------------------------------------------------------------------------------------------------------------------------------

Q. No. 3 Give the name to following inorganic compounds

 (i) H3PO4 (ii) Fe2(SO4)3 (iii) Na2EDTA (iv) XeF6 (v) Na2HPO4 (vi) Na2S2O4

Q. No. 4 Fill in the blanks

 (i) Standard solution is that solution whose concentration is ----------------------

 (ii) When the pH of the solution is more than 7 that solution is ---------------------

1. Structural formula of EDTA is -------------------------------------------------------------------------------------------------------------------------------------
2. Indicators are those which show the ------------------------------ of chemical reaction
3. Acidic radical in salt Na­2SO4 is -------------------------------------------------------
4. Molecular weight of H2C2O4 ----------------------------------------------------------

Q. No. 5 Write the theory procedure observation calculation and the result for the given object by putting imaginary burete readings

 **Object:** Determine the permanent hardness of tape water by titrating with 0.01M solution of EDTA